

(12) INTERNATIONAL APPLICATION PUBLISHED UNDER THE PATENT COOPERATION TREATY (PCT)

(19) World Intellectual Property Organization
International Bureau



(43) International Publication Date
21 March 2002 (21.03.2002)

PCT

(10) International Publication Number
WO 02/22453 A2

(51) International Patent Classification: B65D 39/00

(21) International Application Number: PCT/US01/28378

(22) International Filing Date:
10 September 2001 (10.09.2001)

(25) Filing Language: English

(26) Publication Language: English

(30) Priority Data:
60/232,019 12 September 2000 (12.09.2000) US

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(81) Designated States (*national*): AE, AG, AL, AM, AT, AU,
AZ, BA, BB, BG, BR, BY, BZ, CA, CH, CN, CO, CR, CZ,
DE, DK, DM, DZ, EC, EE, ES, FI, GB, GD, GE, GH, GM,
HR, HU, ID, IL, IN, IS, JP, KE, KG, KR, KZ, LC, LK, LR,
LS, LT, LU, LV, MA, MD, MG, MK, MN, MW, MX, MZ,
NO, NZ, PL, PT, RO, RU, SD, SE, SG, SI, SK, SL, TJ, TM,
TR, TT, TZ, UA, UG, UZ, YU, ZA, ZW.

(84) Designated States (*regional*): ARIPO patent (GH, GM,
KE, LS, MW, MZ, SD, SL, SZ, TZ, UG, ZW), Eurasian
patent (AM, AZ, BY, KG, KZ, MD, RU, TJ, TM), European
patent (AT, BE, CH, CY, DE, DK, ES, FI, FR, GB, GR, IE,
IT, LU, MC, NL, PT, SE, TR), OAPI patent (BF, BJ, CF,
CG, CI, CM, GA, GN, GQ, GW, ML, MR, NE, SN, TD,
TG).

Published:

— without international search report and to be republished
upon receipt of that report

For two-letter codes and other abbreviations, refer to the "Guid-
ance Notes on Codes and Abbreviations" appearing at the begin-
ning of each regular issue of the PCT Gazette.

(54) Title: GAS PERMEATION RESISTANT SYNTHETIC CORK CLOSURE

(57) Abstract: A synthetic cork closure having at least a portion thereof coated with a gas impermeable polymer coating composi-
tion.

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GAS PERMEATION RESISTANT SYNTHETIC CORK CLOSURE

This invention relates to synthetic cork closures for liquid containers and to processes for making such closures.

Synthetic corks are typically made from a foamed polymer and are formed using either a profile extrusion method or injection molding or may be punched out of foamed sheets. Hybrid corks also exist where natural cork is ground and recombined using binding agents. Synthetic corks and processes for preparing them are described in U.S. Patents 5,975,322, 5,904,965, 5,855,287, 5,710,184, 5,496,862 and 4,363,849. All of these synthetic corks and hybrid corks suffer from the uncontrolled permeation of gases in and out of the cork and the scalping of flavors caused by the polymers used.

It would be desirable to provide synthetic corks that have high resistance to the permeation of gases in and out of the cork.

In a first aspect, the present invention is a synthetic cork closure having at least a portion thereof coated with a gas impermeable polymer coating composition.

In a second aspect, the present invention is a process for preparing the closure of the first aspect which comprises providing a synthetic cork closure and coating at least a portion thereof with a gas impermeable polymer coating composition.

In a third aspect, the present invention is a process for preparing the closure of the first aspect which comprises inserting the synthetic closure into a container, dripping the gas impermeable coating composition onto the free end of the closure and allowing the solvent in the coating composition to evaporate.

The synthetic cork closures which can be employed in the practice of the present invention include the synthetic closures described in U.S. Patents 5,975,322, 5,904,965, 5,855,287, 5,710,184, 5,496,862 and 4,363,849. Most preferred synthetic cork closures are those which are commercially available from Neocork Technologies, Inc., Napa Valley, CA.

The gas impermeable coating composition of the present invention can be prepared by dissolving a gas impermeable polymer in a volatile organic solvent solution. Preferably, the organic solvent solution comprises a blend of a solvent and a cosolvent. The organic solvent dissolves the gas impermeable polymer. The cosolvent is employed to

control solvent evaporation rate and to dissolve any additives which can not be dissolved by the solvent. If the gas impermeable polymer is vinylidene chloride polymer, the cosolvent is also used to control the vinylidene chloride polymer crystallinity.

5 The gas impermeable polymers which can be employed in the practice of the present invention include vinylidene chloride polymers and copolymers, polychlorotrifluoroethylene, polyethylene terephthalate and polyisobutylene. The most preferred gas impermeable polymers are vinylidene chloride polymers.

10 The vinylidene chloride polymers which can be employed in the practice of the present invention are well-known in the art. See, for example, U.S. Patents 3,642,743 and 3,879,359. The most common PVDC resins are known as SARAN™ resins, manufactured by The Dow Chemical Company. As used herein, the term "vinylidene chloride polymer" or "PVDC" encompasses homopolymers of vinylidene chloride, and also copolymers and terpolymers thereof, wherein the major component is vinylidene chloride and the remainder is one or more monoethylenically unsaturated monomers copolymerizable
15 with the vinylidene chloride monomer. Monoethylenically unsaturated monomers which can be employed in the practice of the present invention for preparing the vinylidene chloride polymers include vinyl chloride, alkyl acrylates, alkyl methacrylates, acrylic acid, methacrylic acid, itaconic acid, acrylonitrile, and methacrylonitrile. Preferred ethylenically unsaturated monomers include vinyl chloride, acrylonitrile, methacrylonitrile, alkyl
20 acrylates, and alkyl methacrylates. More preferred ethylenically unsaturated monomers include vinyl chloride, acrylonitrile, methacrylonitrile, and the alkyl acrylates and alkyl methacrylates having from 1 to 8 carbon atoms per alkyl group. Most preferred ethylenically unsaturated monomers are vinyl chloride, methyl acrylate, ethyl acrylate, and methyl methacrylate, acrylonitrile, and methacrylonitrile.

25 Preferably, the vinylidene chloride copolymer is (1) a copolymer of (a) from 80 to 93 mole percent vinylidene chloride and (b) from 20 to 7 mole percent of at least one monoethylenically unsaturated monomer copolymerizable therewith (SARAN F-278) or (2) a copolymer of (a) from 65 to 75 mole percent vinylidene chloride and (b) from 35 to 25 mole percent of at least one monoethylenically unsaturated monomer copolymerizable
30 therewith (SARAN F-310).

The above vinylidene chloride polymers and processes for preparing them are well known. See, for example, U.S. Patents 3,817,780; 3,879,359; 4,351,929 and 4,451,632.

The organic solvents which can be employed in the practice of the present invention for preparing the solvent mixture include acetone, methyl ethyl ketone, cyclohexanone, ethyl acetate, n-propyl acetate, isopropyl acetate, propylene oxide, dioxane, tetrahydrofuran and mixtures thereof. Preferred solvents are methyl ethyl ketone, acetone, ethyl acetate and tetrahydrofuran. More preferred solvents are methyl ethyl ketone, acetone, and tetrahydrofuran, with methyl ethyl ketone as the most preferred. Cosolvents which can be employed in the practice of the present invention for preparing the solvent mixture include aliphatic and alicyclic hydrocarbons, such as hexane, heptane, cyclohexane, cyclohexene, methylcyclohexane and dialkylethers, such as t-amyl methyl ether. Most preferred cosolvents are methylcyclohexane and heptane. Most preferred is methylcyclohexane.

A commonly known organic solvent system, which is a mixture of the solvent THF (tetrahydrofuran) and the cosolvent toluene, can be employed in the practice of the present invention for preparing the gas impermeable coating composition.

The coating composition may also contain various additives to impart desirable properties such as, for example, slip properties, to the finished coating.

The amount of solvent and cosolvent used in the solvent mixture depends on many factors such as the composition and solubility of vinylidene chloride polymer, the desired flow characteristics of the coating, desired pot life and drying time, desired coating thickness, and desired wettability of the closure the type of cosolvent employed, vinylidene chloride crystallinity and the temperature at which the vinylidene chloride and other additives are dissolved. In general, the solvent is used in an amount of from 50 to 85 weight percent, preferably from 60 to 80 weight percent and, most preferably, in an amount of from 65 to 75 weight percent, based on the weight of the solvent mixture and the balance is the cosolvent.

The thixotropic agent which can be employed in the practice of the present invention include fumed silica, kaopolite, bentonites, talc, or mixtures thereof.

The amount of thixotropic agent employed in the practice of the present invention for preparing the coating composition depends on desired flow characteristics of the coating, the specific method of applying the coating, and the desired coating thickness. In general, this amount is from 0 to 25 weight percent based on the weight of the composition. Preferably, the gas impermeable polymer coating composition comprises from 5 weight percent to 20 weight percent of a vinylidene chloride polymer, from 70 weight percent to 90 weight percent of an organic solvent or blend of organic solvents and, from 5 weight percent to 10 weight percent of a thixotropic agent. The coating composition can be prepared by methods known in the art such as by dissolving a vinylidene chloride polymer or resin in a volatile organic solvent mixture at a temperature and for a time sufficient to dissolve substantially all of the vinylidene chloride resin.

The synthetic cork closure of the present invention can be prepared by providing a synthetic cork closure and applying a coating of the gas impermeable polymer composition to at least a portion of the surface thereof. For example, the coating may be applied only to the surface of the closure that is likely to contact the contents of the container, such as one or both ends of the synthetic cork closure, or it can be applied to the whole surface of the cork closure.

The coating composition can be applied onto the surface of the synthetic cork closure by methods known in the art such as, for example, by analox gravure coating, offset coating, pad print coating, screen or stencil coating, brush coating, and spray coating. In gravure coating, the coating composition is doctored onto an inert surface and either pressed onto the closure or transferred to an offset pad, which in turn is used to apply the coating to the closure. Likewise, a pad print system can be used to transfer the coating in a controlled amount from a cliché to the closure by means of an elastomeric pad. Brush and spray coating are very conventional techniques which can be used to apply the coating to the closure, albeit in a less-controlled manner. Another method is applying the coating by means of a silk-screen or stencil process involving a squeegee. Depending on the solvent system, for any of these methods, the coating can be applied from an open well or a closed cup system. These are all methods well-known to those skilled in the art of applying coatings to surfaces.

The coating composition can be applied onto the surface of the synthetic cork closure after it has been inserted into a container. For example, the coating composition,

which comprises a vinylidene chloride copolymer dissolved in a solvent, can be applied to the inserted closure by dripping the coating composition onto the free end of the closure and allowing the solvent to evaporate.

Any residual solvent in the coating can be removed by conventional methods
5 such as, for example, by drying the coated substrate in a forced-air oven. Advantageously, the coated substrate is dried for a period of time and at a temperature sufficient to remove the solvent and develop crystallinity, if required, in the finished coating. The temperature and length of time can be easily determined by those skilled in the art without undue experimentation. The following examples are for illustrative purposes only and are not
10 intended to limit the scope of this invention. Unless otherwise indicated, all parts and percentages are by weight.

EXAMPLES

The synthetic corks used in the following Examples were obtained from Neocork Technologies, Inc., Napa Valley, CA.

15 The synthetic corks comprise an outer skin of a styrene ethylene butadiene styrene (SEBS) and an inner core of foamed AFFINITY™ PF1140, a homogeneously branched, substantially linear ethylene/α-olefin polymer, which is a product of The Dow Chemical Company. Some of the synthetic corks were plasma-treated to 42 dynes. The natural corks were purchased from a local wine supply store.

20 The following coating compositions were used in the Examples:

- (1) SARAN™ F-310 and 1 weight percent Solvent Violet 33 dye
- (2) 7.5 weight percent of Coating Composition 1 and 92.5 weight percent methyl ethyl ketone (MEK)
- (3) 15 weight percent of Coating Composition 1 and 85 weight percent methyl ethyl
25 ketone (MEK)
- (4) Polyurethane Primer - 9 parts by weight MORTON ADCOTE™ 57R2 and 1 part by weight COREACTANT™ F
- (5) A mixture of 10 percent by weight SARAN F-310, 8 percent by weight KORTHIX™, 8 percent toluene, and 74 percent MEK

- (6) A mixture of 25 percent by weight SARAN F-310 and 75 percent by weight of acetone.

SARAN F-310 is a copolymer of from 65 to 75 mole percent vinylidene chloride and (b) from 35 to 25 mole percent of at least one monoethylenically unsaturated monomer copolymerizable therewith.

Solvent Violet 33 dye is an anthraquinone dyestuff available from Mitsubishi Chemical Corporation as Diaresin™ Blue J. Solvent Violet 33 has a melting point of 143°C.

The Solvent Violet 33 dye was added to the solution so that the coatings could be seen.

ADCOTE™ 57R2 is a polyester available from Morton International.

COREACTANT™ F is an isocyanate available from Morton International.

Polyurethane Primer - nine parts by weight of ADCOTE 57R2 blended with 1 part of COREACTANT F, then diluted with an equal amount of methyl ethyl ketone to achieve the desired viscosity.

KORTHIX™ is a bentonite clay available from Kaopolite Corporation.

Example 1 - Coverage and Adhesion Tests

Coating compositions 2 and 3 were applied to some of the corks by painting, using a paint brush, rolling, using a paint roller, dipping and dripping onto the surface of the cork. The coated corks were allowed to dry in a convection air oven set at 70°C for 5 minutes. The coated and uncoated corks were then examined using a light microscope to look at the coverage of the barrier coating.

The violet solution was found to completely cover the surface of the cork. Using light microscopy the violet solution was seen to have penetrated down into the open cells of the cork. A sharp pick was used to try and "pick-off" the coating on the cork. Samples of the cork were crushed using a hand held, jaw type corking unit. It was found that the coating could not be removed from both the treated and untreated cork surfaces. Adhesive tape also failed to remove the coating on the cork.

Example 2 - Natural Corks

Both ends of 20 natural corks were coated with either Coating Composition 1 or 2. No pretreatment to increase the adhesion between the cork surface and the coating composition was done. Excellent adhesion was observed between the cork and the barrier coating.

Example 3- Bottle Trials

The following corks were inserted into 750 mL wine bottles. The corks were coated with Coating Composition 1 or 2, as indicated in Table 1. The corks were also spray-coated with a silicone lubricant to facilitate easy removal from the bottle. Samples 1S-9S are synthetic corks and Samples 1N to 7N are natural corks.

Table 1

Cork Type	Coating Composition	# of Corks Coated	Surface Coated	percent PVDC Solids
1S	2	10	Both ends	7 percent
2S	3	10	Both ends	15 percent
3S	2	5	Both ends	7 percent
4S	3	5	Both ends	15 percent
5S	2	5	All surfaces	7 percent
6S	3	5	All surfaces	15 percent
7S	2	5	All surfaces	7 percent
8S	3	5	All surfaces	15 percent
9S	-	10	None	0 percent
1N	2	5	Both ends	7 percent
2N	3	5	Both ends	15 percent
3N	2	5	All surfaces	7 percent
4N	3	5	All surfaces	15 percent
5N	2	5	All surfaces	7 percent
6N	3	5	All surfaces	15 percent
7N	-	5	None	0 percent

Example 4

The corks described in Table 1 were inserted into wine bottles. The tops of the bottles were cut off and mounted on a plate using epoxy. The gas permeation rates were determined using a Mocon test unit. The following Tables 2 and 3 show that application of the SARAN coating to the natural cork significantly reduced the permeation property of the

5 the natural cork and reduced the permeation properties of the synthetic cork by more than 50 percent.

Table 2

SAMPLES A, B, C, D AND E ARE NON-COATED SAMPLES

Comparative Example	A	B	C	D	E
Cork Type	7N	7N	7N	9S	9S
#days purge					
time in Vacuum	0	4	0	0	7
	OTR*	OTR*	OTR*	OTR*	OTR*
Day 1	12.91	12.02		too high	
Day 2	12.65	10.08		0.1792	
Day 3	11.59	9.95		0.1425	
Day 4	12.73	11.04		0.1314	0.1352
Day 5	11.07	10.05		0.1157	0.0990
Day 6		9.42	13.02	0.1117	0.0974
Day 7		9.76	12.79	0.1033	0.0801
Day 8		9.25	12.25	0.0985	0.0762
Day 9		9.74	12.25	0.0895	0.0690
Day 10		12.02		0.0894	0.0682
Day 11		10.08		0.0847	0.0657
Day 12		9.95		0.0848	0.0652
Day 13		11.04		0.0828	0.0638
Day 14		10.05		0.0773	0.0582
Day 15		9.76		0.0743	0.0656
Day 16				0.0743	0.0605
Day 17		9.25		0.0704	0.0585
Day 18		9.74		0.0671	0.0562
Day 19				0.0683	0.0613
	Epoxy applied to edge of Cork #1				
		8.38			
Final OTR*	11.07	8.38	12.25	0.0683	0.0613

*OTR - Oxygen transmission rate in cc/bottle-day-atm.

Table 3
Table 3 Saran Coated Samples - Both ends coated

Example	1	2	3	4	5	6	7	8
Cork Type	1S	1S	2S	2S	1N	1N	2N	2S
percent PVDC solids	7.50 percent	7.50 percent	15.00 percent	15.00 percent	7.50 percent	7.50 percent	15.00 percent	15.00 percent
	OTR*	OTR*	OTR*		OTR*	OTR*	OTR*	OTR*
Day 1	0.0933		0.0915		1.6102		1.2564	
Day 2	0.0382		0.0386		2.2332		1.2234	
Day 3	0.0301		0.0305		2.2044		1.3698	
Day 4	0.0289		0.0293		2.2240		1.353	
Day 5	0.0262		0.0258		2.2220		1.194	
Day 6	0.0251		0.0248		2.2184		1.258	
Day 7	0.0228		0.023		2.0976		1.1932	
Day 8	0.0223		0.0219		1.9648		1.6856	
Day 9	0.0208		0.0207		1.9150		1.3365	
Day 10	0.0212		0.0211		1.9560		1.4543	
Day 11	0.0204		0.0206		1.9662		1.3652	
Day 12	new samples loaded							
Day 13		0.0828		0.0638		3.56		2.99
Day 14		0.0415		0.0418		2.05		2.6
Day 15		0.0375		0.0373		3.48		2.64
Day 16		0.0343		0.0340		3.55		2.68
Day 17		0.3150		0.0314		3.59		2.6
Day 18		0.0294		0.0300		3.65		2.68
Day 19		0.0282		0.0282		3.77		2.81
Final OTR	0.020	0.028	0.021	0.028	1.966	3.770	1.365	2.810

*OTR - Oxygen transmission rate in cc-cc/bottle-day-atm.

Example 5

The coating of 1 vs. 2 sides of the synthetic cork were compared. No difference was seen between coating one end of the cork versus two ends, as indicated in

5 Table 4.

Table 4
Oxygen Transmission Rate* – Steady State*

Sample	1 Day	5 Days
Synthetic - not coated	0.18	0.12
Synthetic - one end coated	0.09	0.04
Synthetic - both ends coated	0.09	0.04

*Oxygen transmission rate in cc/bottle-day-atm.

10 Example 6

Synthetic corks without the coating composition were inserted into bottles to a sufficient depth such that the top of the cork was recessed two millimeters below the top of the bottle neck. Composition number 6 was then poured onto the exposed top of the cork in a quantity sufficient to fill the volume between the top of the cork and the top of the
15 bottle such that the surface of the dried composition was nearly flush with the top of the bottle neck. After drying, it was observed that the dried composition was adherent to the glass of the bottle and to the cork, providing a permeation-resistant seal. The dried composition was sufficiently flexible to allow penetration by a conventional corkscrew for removal of the cork.

20

WHAT IS CLAIMED IS:

1. A synthetic cork closure having at least a portion thereof coated with a gas impermeable polymer.
2. The synthetic cork closure of Claim 1 wherein only one end of the closure
5 is coated with the gas impermeable polymer.
3. The synthetic cork closure of Claim 1 wherein both ends of the closure are coated with the gas impermeable polymer.
4. The synthetic cork closure of Claim 1 wherein the entire surface of the closure is coated with the gas impermeable polymer.
- 10 5. The synthetic cork closure of Claim 1 wherein the gas impermeable polymer is a vinylidene chloride polymer.
6. The synthetic cork closure of Claim 5 wherein the vinylidene chloride polymer is (1) a copolymer of (a) from 80 to 93 mole percent vinylidene chloride and (b) from 20 to 7 mole percent of at least one monoethylenically unsaturated monomer
15 copolymerizable therewith or (2) a copolymer of (a) from 65 to 75 mole percent vinylidene chloride and (b) from 35 to 25 mole percent of at least one monoethylenically unsaturated monomer copolymerizable therewith.
7. A process for preparing the synthetic cork closure of Claim 1 which comprises providing a synthetic cork closure and coating at least a portion thereof with a gas
20 impermeable polymer.
8. The process of Claim 7 wherein the gas impermeable polymer is a vinylidene chloride polymer.
9. The process of Claim 8 wherein the vinylidene chloride polymer is (1) a copolymer of (a) from 80 to 93 mole percent vinylidene chloride and (b) from 20 to 7 mole
25 percent of at least one monoethylenically unsaturated monomer copolymerizable therewith or (2) a copolymer of (a) from 65 to 75 mole percent vinylidene chloride and (b) from 35 to 25 mole percent of at least one monoethylenically unsaturated monomer copolymerizable therewith.
10. A process for preparing the synthetic cork of Claim 1 which comprises
30 inserting a synthetic cork closure into a container, applying onto the free end of the synthetic

cork closure a coating composition comprising a vinylidene chloride polymer dissolved in a solvent and allowing the solvent to evaporate.

11. A gas impermeable polymer coating composition comprising from 5 weight percent to 20 weight percent of a vinylidene chloride polymer, from 70 weight percent to 90 weight percent of an organic solvent or blend of organic solvents and from 5 weight percent to 10 weight percent of a thixotropic agent.

12. The gas impermeable polymer coating composition of Claim 11 wherein the gas impermeable polymer is a vinylidene chloride polymer.

13. The gas impermeable polymer coating composition of Claim 12 wherein the vinylidene chloride polymer is (1) a copolymer of (a) from 80 to 93 mole percent vinylidene chloride and (b) from 20 to 7 mole percent of at least one monoethylenically unsaturated monomer copolymerizable therewith or (2) a copolymer of (a) from 65 to 75 mole percent vinylidene chloride and (b) from 35 to 25 mole percent of at least one monoethylenically unsaturated monomer copolymerizable therewith.

14. The gas impermeable polymer coating composition of Claim 11 wherein the organic solvent is selected from acetone, methyl ethyl ketone, cyclohexanone, ethyl acetate, n-propyl acetate, isopropyl acetate, dibutyl ether, propylene oxide, dioxane, toluene and tetrahydrofuran.

15. The gas impermeable coating composition of Claim 11 wherein the thixotropic agent is selected from fumed silica, kaopolite, bentonite, talc and mixtures thereof.